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Chapter 6 Solutions Engineering and Chemical Thermodynamics Wyatt Tenhaeff Milo Koretsky Department of Chemical Engineering Oregon State University 2 6.1 (a) The Clausius-Clapeyron equation: dP i sat P i sat = ! h i vap dT RT 2 or ln P i sat 101 kPa [ ] " # $ % & ' = ( ) h i vap R 1 T (1 373 [K] * + , - . / so P i sat = 101 kPa [ ] ( ) exp " # h i vap R 1 T " 1 373 [K] $ % & ' ( ) * + , - .

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Chapter 6 2 If the fluid is in thermodynamic equilibrium any thermodynamic variable for a pure substance, like pure water, can be written in terms of any two other thermodynamic variables , i.e. $p=p(\rho,T)$ (6.1.1) where the functional relationship in depends on the substance.

Chapter 6 Thermodynamics and the Equations of Motion

Chapter 6: Solution Thermodynamics and Principles of Phase Equilibria In all the preceding chapters we have focused primarily on thermodynamic systems comprising pure substances. However, in all of nature, mixtures are ubiquitous.

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