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(aq) $32 \text{ Ag (aq) NO}_3 \text{ (aq) O K (aq) NO}_3$
(aq) AgI(s)

Chemical Reactions

In 30 minutes, $5 \times 10^5 \text{ J s}^{-1} \text{ m}^{-2} \times 1800$
 s gives $9 \times 10^8 \text{ J m}^{-2}$, which
corresponds to 1.36×10^{27} molecules
 m^{-2} , or 1.36×10^{23} molecules cm^{-2} .
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is -73 J and in the second case $q = -3 \text{ J}$.
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