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the following questions in the space
provided. 1. Given the following
equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O c. What is the mole ratio of O_2 to H

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Chapter 9 - Stoichiometry 9-1

Introduction to Stoichiometry

Composition Stoichiometry - deals with mass relationships of elements in compounds
Reaction Stoichiometry - Involves mass relationships between reactants and products in a chemical reaction
I. Reaction Stoichiometry Problems A.

Chapter 9 - Stoichiometry

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Modern Chemistry 77 Stoichiometry
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SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H

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SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many grams of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. Given the following equation ...

CHAPTER 9 REVIEW

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students...

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Show all your work in the space

provided. 1. The following equation

represents a laboratory preparation for

oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) +$

$3\text{O}_2(\text{g})$ How many moles of O_2 form if

3.0 mol of KClO_3 are totally consumed?

2. Given the following equation: $\text{H}_2(\text{g}) +$

$\text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

CHAPTER 9 REVIEW

CHAPTER 9 REVIEW. Stoichiometry.

SECTION 1. SHORT ANSWER Answer the

following questions in the space

provided. 1. _____ The coefficients in a

chemical equation represent the.

(a) masses in grams of all reactants and

products. (b) relative number of moles of

reactants and products. (c) number of

atoms of each element in each

compound in a reaction.

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in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N_2 are mixed with 120 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g)$

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