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8.4 $-5-17 -22 = x10 = 2.05 \times 10$. 17.
4.1 $\times 10$. 4.1. b) 5.4 $\times 10$. 7.3 $\times 10$. 32. 14 =
5.4. 7.3. $\times 10$. 32-14 = .74 $\times 10$. 18 =
7.4 $\times 10$. 17. Information: Adding and
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3. Iodine has 17 more electrons than krypton and so you can begin by writing [Kr] followed by orbitals for 617 more electrons. After 4p comes 5s²... Iodine =

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[Kr]5s²4d¹⁰5p⁵ Any of the noble gases can be used for abbreviations. Here are a few more examples: iron = [Ar]4s²3d⁶ cesium = [Xe]6s¹ phosphorus = [Ne]3s²3p³

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Answer Key Converting Between Grams

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Given your answer to number 7, you should be able to answer the following questions. a) Find the pH of a solution that has a pOH of 4.6. $9.4; \text{pH} + \text{pOH} =$

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14 $\text{pH} = 14 - \text{pOH} = 9.4$ b) The $[\text{H}_3\text{O}^+]$ of a solution is $2.55 \times 10^{-3} \text{ M}$.

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7; since chlorine is in column 17 (or "7A") it has 7 valence electrons. 3. Since CCl_4 is made up of one carbon and four chlorine atoms, how many total valence

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electrons does CCl_4 have? Hint: add your answer to question 1 and four times your answer to question 2. $4 + 4(7) = 4 + 28 = 32$. Verify that there are 32 electrons pictured in the ...

ChemQuest 23 : Drawing Covalent Compounds

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Key Chemquest 49 Answer Key The Online Books Page: Maintained by the University of Pennsylvania, this page lists over one ... Iodine is 80% ^{127}I 17% ^{126}I , and 3% ^{128}I . Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that

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24 ChemQuest 8 Name: _____ Date:
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of the Atom Note the following symbols:
(they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6 C 6C (6 ...

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24 ChemQuest 8 Name:

Calculate the average atomic masses. Round all answers to two decimal places. Record your answer with the correct units. 1. Iodine is 80% 127.1, 17% 126.1, and 3% 128.1. Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of

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lithium, which occurs as two isotopes that have the following

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118 ChemQuest 39 Name: _____ Date:
_____ Hour: _____ Information: Molarity
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the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

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