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SAHOTA ELECTROCHEMISTRY STUDY GUIDE- Written - Page 7 of 21 U 9 42. b)

Calculate the initial voltage for the electrochemical cell in part a). (1 mark)

c) Which two metals from the above list would produce an electrochemical cell with the smallest initial voltage? (1 mark)

U 9 43. U 9 44. Consider the following electrochemical cell:

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oxidation reduction ... Galvanic Cells (Voltaic Cells) All about Galvanic Cells, which are also called Voltaic Cells. These are devices that use a chemical reaction to create electricity. Electrolysis Electrolysis is a process where you use electrical energy

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ELECTROCHEMISTRY Notes E

Electrochemistry is the study of reactions in which charged particles (ions or electrons) cross the interface between two phases of matter, typically a metallic phase (the electrode) and a conductive

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Electrochemistry: the study of the interchange of chemical and electrical energy
Redox reaction: a transfer of electrons from the reducing agent to the oxidizing agent
Oxidation: a loss of electrons (an increase in the oxidation number)
Reduction: a gain of electrons (a decrease in the oxidation number)

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processes utilize oxidation and reduction reactions.

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Oxidizing agent and reducing agent
Corrosion of a metal Oxidation number
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