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Chapter 9

average atomic mass of Cl = contribution of Cl + contribution of Cl17 17 = 34.97 u × 0.7578 + 36.97 u × 0.2422 = 26.5003 u + 8.9541 u = 35.4544 u = 35.45 u. The average atomic mass of chlorine is 35.45 u. Chemistry 11 Solutions. 978-0-07-105107-1 Chapter 1 Elements and the Periodic Table • MHR | 2.

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chemistry 11 Check Your Solution The mole ratio of the reactants Na2S2O3 to Cl2 to H2O from the balanced equation is 1:4:5. The calculated mole ratio from the given masses of reactants was 853.778:705.22: 13207.55, which is 1.2:1:18.7 in its lowest ratio (dividing all three by 705.22).

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